

### Configuration Writing Practice

A. Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

1. Na:  $1s^2, 2s^2, 2p^6, 3s^1$
2. Pb:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14}$
3. Sr:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$
4. U:  $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2, 3d^{10}, 4p^6, 5s^2, 4d^{10}, 5p^6, 6s^2, 4f^{14}, 5d^{10}, 6p^6, 7s^2, 5f^3$
5. N:  $1s^2, 2s^2, 2p^3$

B. Write a ground state electron configuration for these ions. Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example:  $N^{3-}$  is  $1s^2 2s^2 2p^6$ . It has three extra electrons

B. Write a ground state electron configuration for these ions. Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example:  $N^{3-}$  is  $1s^2 2s^2 2p^6$ . It has three extra electrons

11.  $O^{2-}$ :  $1s^2, 2s^2, 2p^6$
12.  $Fe^{2+}$ :  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
13.  $B^{3+}$ :  $1s^2$
14.  $Ni^{2+}$ :  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^8$
15.  $K^+$ :  $1s^2 2s^2 2p^6 3s^2 3p^6$
16.  $Co^{3+}$ :  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$

C. For the following electron configurations determine the possible elements (or ions) they may represent

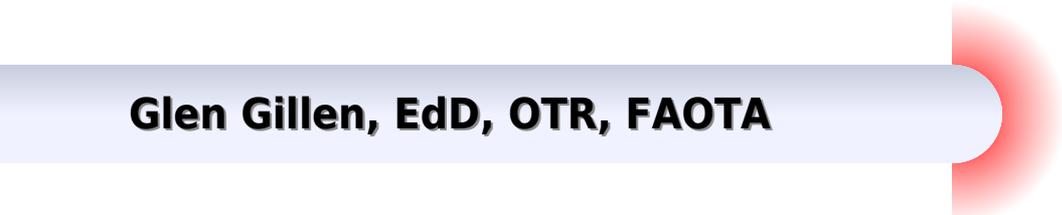
17.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$  : Se
18.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$  : Kr
19.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{10}$  : Dy
20.  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^5$  : Mn
21.  $[Kr] 5s^2 4d^{10} 5p^3$ : As
22.  $[Ar] 4s^1$  : K
23.  $[Xe] 6s^2 4f^{10}$  : Dy
24.  $[Ne] 3s^2 3p^1$  : Al

C. In the space below, write the full (unabbreviated) electron configurations of the following elements:

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$
- 5) neptunium  $1s^2, 2s^2, 2p^6, 3s^2, 3p^6, 4s^2, 3d^{10}, 4p^6, 5s^2, 4d^{10}, 5p^6, 6s^2, 4f^{14}, 5d^{10}, 6p^6, 7s^2, 5f^6$

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**Glen Gillen, EdD, OTR, FAOTA**



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main venues for discussing recent trends and issues This volume contains refereed papers and reports from the 1995 meeting The materials cover a broad range of HCI related topics including visualization computer supported communication task analysis formal methods user support and cyberspace The documents consider both research and commercial perspectives making the book essential for all researchers designers and manufacturers who need to keep abreast of developments in HCI

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**Chemistry** James E. Brady, Fred Senese, 2004-02-04 Publisher Description

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**Fundamentals of Chemistry** David E. Goldberg, 2001

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*Mandarin Chinese Dual Language Immersion Programs* Ko-Yin Sung, Hsiao-Mei Tsai, 2019-06-05 This book discusses multiple aspects of Chinese dual language immersion DLI programs with a focus on the controversial Utah model The first

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