

# Chm 113 Final Study Guide Questions & Answers 100% Correct.

What is the  $[H^+]$  of a solution with a pH of 3.78? - ANSWER  
 $pH = 3.78$   
 $[H^+] = 10^{-3.78}$   
 $= 1.66 \times 10^{-4}$

If HX represents the formula of a weak acid, what is the formula of the conjugate base of HX? - ANSWER  
If HX weak acid then the formula of the conjugate base of X-

Let HX represent the formula of a weak acid. Write a balanced chemical equation for the reaction of this weak acid with the strong base NaOH? You may omit any spectator ions if you wish. - ANSWER  
 $HX + NaOH \rightarrow NaX + H_2O$   
For the equation you wrote above, what two terms are considered/assumed equal? - ANSWER  
 $[HA] = [A^-]$

Regarding today's experiment, what is the weak acid we are going to titrate and which strong base is used to titrate this weak acid? - ANSWER  
Weak Acid: Acetic Acid  
Strong Base: NaOH

If a weak acid, HX, is titrated with a strong base such as NaOH, will the equivalence point in the titration occur below pH 7.0, at pH 7.0 or above pH 7.0? - ANSWER  
the equivalence point in the titration will be somewhat above pH 7.0

A student is presented with a 0.65 M solution of an unknown weak acid. The student measures the pH of that solution and determines a value of 3.78 for the unknown solution. What, then, is the  $K_a$  of this weak acid? - ANSWER  
 $K_a = \frac{(1.66 \times 10^{-4})^2}{.65 - 1.66 \times 10^{-4}} = 4.2 \times 10^{-8}$

What is the acid dissociation constant ( $K_a$ )? - ANSWER  
 $K_a = \frac{[H_3O^+][A^-]}{[HA]}$

What is the pH of a weak acid whose  $K_a$  is  $6.03 \times 10^{-6}$ , and the ratio of  $[A^-]/[HA]$  is 3. - ANSWER  
 $6.03 \times 10^{-6} = \frac{[9]/3}{[HA]}$   
 $pH = 5.69$

# 113 Study Guide Answers 129346

**Camilla Rothe**



**113 Study Guide Answers 129346:**

## **113 Study Guide Answers 129346** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**113 Study Guide Answers 129346**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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