

**CHEMISTRY 103 – Practice Sheet #4 – Naming Compounds and Writing Formulas
Chapter 2**

Prepared by Dr. Tony Jacob

<http://www.chem.utsc.edu/aneas/cic/> (Resource page)

1. Name the compounds below. Includes molecular and ionic compounds, polyatomic ions, acids, and transition metals.

a. Na_2S sodium sulfide

b. CaO calcium oxide

c. N_2O dinitrogen monoxide

d. K_3N potassium nitride

e. NO nitrogen oxide

f. Li_3P lithium phosphide

g. NaNO_3 sodium nitrate

h. $\text{Ca}_3(\text{PO}_4)_2$ calcium phosphate

i. MgF_2 magnesium fluoride

j. $(\text{NH}_4)_2\text{SO}_4$ ammonium sulfate

k. P_2O_3 diphosphorous pentoxide

l. NaOH sodium hydroxide

m. HBr hydrogen bromide

n. $\text{Sc}(\text{OH})_3$ scandium III hydroxide

o. HNO_3 nitric acid

p. NH_4NO_3 ammonium nitrate

q. FeCl_3 ferric chloride

r. $(\text{NH}_4)_2\text{O}$ ammonium oxide

s. CuO copper oxide

t. HCl hydrochloric acid

u. NaHCO_3 sodium bicarbonate

v. $\text{Al}(\text{C}_2\text{H}_3\text{O}_2)_3$ aluminum acetate

w. H_2SO_4 dihydrogen sulfate

x. HI hydroiodic acid

y. SO_3 sulfur trioxide

z. $\text{HC}_2\text{H}_3\text{O}_2$ acetic acid

aa. $\text{Ca}(\text{OH})_2$ calcium hydroxide

bb. $\text{Mn}(\text{ClO}_4)_3$ Manganese III perchlorate

cc. HClO_4 perchlorate acid

dd. $\text{Zn}(\text{CN})_2$ zinc II cyanide

ee. $\text{NaC}_2\text{H}_3\text{O}_2$ sodium acetate

ff. K_2SO_4 potassium sulfate

gg. H_2CO_3 dihydrogen carbonate

hh. H_3PO_4 trihydrogen phosphate

ii. HF hydrofluoric acid

jj. CH_3COOH acetic acid

kk. LiOH lithium hydroxide

ll. H_3O^+ trihydroxide ion

mm. NH_3 ammonia

nn. CH_4 methane

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