

# General Chemistry cheat sheet

## DAT General Chemistry Equation Sheet

### Chapter 0: General and Lab Concepts Review

<b>Dilutions</b>	$M_1V_1 = M_2V_2$ or $C_1V_1 = C_2V_2$	$M$ or $C$ = concentration $V$ = volume
<b>Percent Error</b>	$\frac{(A - T)}{T} \times 100$	$T$ = theoretical $A$ = actual
<b>Absorbance (Spectrophotometer)</b>	$Abs = \epsilon cl$	$\epsilon$ = molar extinction coefficient (molar absorptivity) $c$ = sample's concentration $l$ = path length

### Chapter 2: Atomic and Electronic Structure

<b>Energy of a photon</b>	$E_{\text{photon}} = hf = \frac{hc}{\lambda}$	$h$ = Planck's constant ( $6.63 \times 10^{-34} \text{ J} \cdot \text{s}$ ) $f$ = photon's frequency $c$ = speed of light ( $3.0 \times 10^8 \text{ m/s}$ ) $\lambda$ = photon's wavelength
<b>Absorption/Emission Line Spectra</b>	$\Delta E = E_{\text{photon}}$	
<b>Kinetic Energy of an electron (Photo-electric Effect)</b>	$KE_{e^-} = E_{\text{photon}} - \phi$	$\phi$ = work function (minimum energy needed to ionize electron)

### Chapter 7: Chemical Solutions

<b>Molarity</b>	$M = \frac{\text{moles}_{\text{solute}}}{L_{\text{solution}}}$	
<b>Molality</b>	$m = \frac{\text{moles}_{\text{solute}}}{kg_{\text{solvent}}}$	
<b>Henry's Law</b>	$P_A = k_H[A]$	$P_A$ = partial pressure of gas A $k_H$ = Henry's Law constant (varies per problem) $[A]$ = conc. of gas A $i$ = van't Hoff factor $K_p$ = $P, P$ , depression constant $m$ = molality
<b>Freezing Point Depression</b>	$\Delta T_F = -iK_F m$	
<b>Boiling Point Elevation</b>	$\Delta T_B = iK_B m$	$i$ = van't Hoff factor $K_B$ = $B, P$ , depression constant $m$ = molality
<b>Vapor Pressure Depression (Raoult's Law)</b>	$P_{\text{solute}} = X_{\text{solute}} P_{\text{solute}}^0$	$P_{\text{solute}}$ = VP of solution $X_{\text{solute}}$ = mol fract of solvent $P_{\text{solute}}^0$ = VP of pure solvent
<b>Osmotic Pressure (<math>\pi</math>)</b>	$\pi = iMRT$	$M$ = molarity of solute $i$ = van't Hoff factor $R$ = $0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$ $T$ = temp. in Kelvins

### Chapter 5: Gases

<b>Pressure</b>	$P = \frac{F}{A}$	$F$ = force $A$ = area
<b>Average Kinetic Energy</b>	$KE_{\text{avg}} = \frac{3}{2} RT$	$R = 0.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$
<b>Root-Mean-Square Speed (<math>v</math>)</b>	$v = \sqrt{\frac{3RT}{M_m}}$	$R = 0.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$ $M_m$ = molar mass
<b>Ideal Gas Law</b>	$PV = nRT$	$n$ = # of moles $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$
<b>Boyle's Law</b>	$V \propto \frac{1}{P}$	
<b>Charles' Law</b>	$V \propto T$	
<b>Avogadro's Law</b>	$V \propto n$	
<b>Combined Gas Law</b>	$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$	
<b>Standard Temp. &amp; Pressure (STP)</b>	$P = 1 \text{ atm}$ $T = 273 \text{ K}$	*1 mol of gas = 22.4 L at STP
<b>Standard Conditions</b>	All aqueous species @ 1M All gaseous species @ 1 atm $T = 298 \text{ K}$	
<b>Density</b>	$\frac{P(MM)}{RT} = \frac{m}{V}$	$MM$ = molar mass $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$ $m$ = mass $V$ = volume
<b>Dalton's Law of Partial Pressures</b>	$P_{\text{total}} = P_A + P_B + \dots$	
<b>Dalton's Law of Partial Pressures</b>	$P_A = X_A P_{\text{total}}$	$X_A$ = mol fraction of gas A
<b>Graham's Law of Effusion</b>	$\frac{r_1}{r_2} = \sqrt{\frac{M_{m2}}{M_{m1}}}$	$r$ = rate of effusion $M$ = molar mass
<b>Real Gas Equation</b>	$(P + \frac{an^2}{V^2})(V - nb) = nRT$	$a$ & $b$ = constants specific to each gas $\frac{an^2}{V^2}$ corrects for IMFs $-nb$ corrects for volume

### Chapter 8: Chemical Kinetics

<b>General</b>	$A + B \rightarrow C + D$	$k$ = rate constant $m$ & $n$ = determined experimentally
<b>Rate Law</b>	$\text{rate} = k[A]^m[B]^n$	$k$ = rate constant $M$ = molarity $s$ = seconds
<b>Rate Constant</b>	0 order: $k = M^1 \cdot s^{-1}$ 1 <sup>st</sup> order: $k = s^{-1}$ 2 <sup>nd</sup> order: $k = M^{-2} \cdot s^{-1}$ 3 <sup>rd</sup> order: $k = M^{-3} \cdot s^{-1}$	
<b>Units</b>		
<b>Arrhenius Equation</b>	$k = Ae^{-E_a/RT}$	$k$ = rate constant $A$ = unique to each rxn $E_a$ = act. energy $R = 0.314 \frac{\text{J}}{\text{mol} \cdot \text{K}}$ $T$ = temp. in Kelvins

# Chemistry Midterm Note Cheat Sheet

**Raffaella Di Napoli**



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