



CHEMISTRY CHEAT SHEET



EQUILIBRIUM

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Chemical Equilibrium occurs in a reversible reaction when the rate of the forward reaction equals the rate of the reverse reaction. Equilibrium constants are represented by the symbol K_c .

What is Equilibrium?
 A reversible reaction is one that can proceed in either direction. At equilibrium, the rate of the forward reaction is equal to the rate of the reverse reaction. The concentrations of the reactants and products remain constant over time.

Equilibrium Constant K_c
 The equilibrium constant K_c is a ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of their stoichiometric coefficient. For a general reaction:



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Reaction	Equilibrium Constant K_c
$aA + bB \rightleftharpoons cC + dD$	$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
$aA + bB \rightleftharpoons cC + dD + eE$	$K_c = \frac{[C]^c [D]^d [E]^e}{[A]^a [B]^b}$
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Factor	Effect on Equilibrium	Example
Concentration	<ul style="list-style-type: none"> Increasing the concentration of a reactant or product will shift the equilibrium position to favor the formation of the products or reactants, respectively. 	$2\text{NO}_2(g) \rightleftharpoons \text{N}_2\text{O}_4(g)$ Adding more NO_2 will shift the equilibrium to the right, favoring the formation of N_2O_4 .
Pressure	<ul style="list-style-type: none"> Increasing the pressure will shift the equilibrium position to favor the formation of the products with the smaller number of moles of gas. 	$\text{N}_2(g) + 3\text{H}_2(g) \rightleftharpoons 2\text{NH}_3(g)$ Increasing the pressure will shift the equilibrium to the right, favoring the formation of NH_3 .
Temperature	<ul style="list-style-type: none"> Increasing the temperature will shift the equilibrium position to favor the formation of the products if the reaction is endothermic, and to favor the formation of the reactants if the reaction is exothermic. 	$\text{N}_2(g) + \text{O}_2(g) \rightleftharpoons 2\text{NO}(g)$ This reaction is endothermic. Increasing the temperature will shift the equilibrium to the right, favoring the formation of NO .

Chemistry Study Guide Answers Chemical Equilibrium

Morris Hein, Susan Arena



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