

Worksheet #3—Stoichiometry

Directions: Solve the following stoichiometry problems. Assume all gases are at STP. Don't forget to write out your balanced chemical equation. Make sure to show ALL work to receive FULL CREDIT!

1. Calculate the number of moles of oxygen required to react completely with 14.8 mol Al in the following equation: $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$

$$\frac{14.8 \text{ moles Al}}{4 \text{ mole Al}} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole Al}} = \underline{11.1 \text{ mol O}_2}$$

3. If 20.0g grams of zinc react with excess hydrochloric acid, how many grams of zinc chloride (ZnCl_2) are produced?
 $\text{Zn}_{(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{ZnCl}_{2(aq)} + \text{H}_{2(g)}$

$$\frac{20 \text{ g Zn}}{65 \text{ g Zn}} \times \frac{1 \text{ mol Zn}}{1 \text{ mol Zn}} \times \frac{1 \text{ mole ZnCl}_2}{1 \text{ mole Zn}} \times \frac{136 \text{ g ZnCl}_2}{1 \text{ mole ZnCl}_2} = \underline{41.8 \text{ g ZnCl}_2}$$

5. How many grams of copper are required to replace 4.00 grams of silver nitrate which are dissolved in water?
 $\text{Cu}_{(s)} + 2\text{AgNO}_{3(aq)} \rightarrow \text{Cu(NO}_3)_2(aq) + 2\text{Ag}_{(s)}$

$$\frac{4.0 \text{ g AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{1 \text{ mol AgNO}_3}{170 \text{ g AgNO}_3} \times \frac{2 \text{ mol AgNO}_3}{1 \text{ mol Cu}} \times \frac{63.5 \text{ g Cu}}{1 \text{ mole Cu}} = \underline{3.0 \text{ g Cu}}$$

2. Calculate the moles of oxygen produced from the decomposition of 0.61 moles of potassium chlorate (KClO_3).



$$\frac{0.61 \text{ moles KClO}_3}{2 \text{ mole KClO}_3} \times \frac{3 \text{ moles of O}_2}{1 \text{ mole KClO}_3} = \underline{0.915 \text{ mol O}_2}$$

4. How many grams of chlorine gas must be reacted with excess sodium iodide (NaI) if 10.0 grams of sodium chloride (NaCl) are needed?
 $2\text{NaI}_{(aq)} + \text{Cl}_{2(g)} \rightarrow 2\text{NaCl}_{(aq)} + \text{I}_{2(g)}$

$$\frac{10 \text{ g NaCl}}{58.5 \text{ g NaCl}} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} \times \frac{2 \text{ mol NaCl}}{1 \text{ mole Cl}_2} \times \frac{71 \text{ g Cl}_2}{1 \text{ mole Cl}_2} = \underline{6.07 \text{ g Cl}_2}$$

6. How many moles of hydrogen gas are needed to react with 15.1 g chlorine gas to produce hydrogen chloride gas?
 $\text{H}_{2(g)} + \text{Cl}_{2(g)} \rightarrow 2\text{HCl}_{(g)}$

$$\frac{15.1 \text{ g Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mole Cl}_2}{71 \text{ g Cl}_2} \times \frac{1 \text{ mole H}_2}{1 \text{ mole Cl}_2} = \underline{0.21 \text{ moles H}_2}$$

Chemistry Stoichiometry Problem Sheet 2 Answers

Mary Millhollon, Richard H. Langley



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