

**Balancing a Redox Equation for an Acidic Solution  
Using the Oxidation Number Method  
(Student textbook page 615)**

35. Use the oxidation number method to balance the following equation for the combustion of carbon disulfide:



**What Is Required?**

You must balance an equation for the combustion of carbon disulfide, using the oxidation number method.

**What Is Given?**

You are given the reactants, carbon disulfide and oxygen, and the products, carbon dioxide and sulfur dioxide.

You are given rules for assigning oxidation numbers in **Table 9.3** on page 604 of the student textbook.

Plan Your Strategy	Act on Your Strategy
<p><b>Step 1</b> Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction.</p>	<p><math>\overset{+4}{\text{C}}\overset{-2}{\text{S}}_2(\text{g}) + \overset{0}{\text{O}}_2(\text{g}) \rightarrow \overset{+4}{\text{C}}\overset{-2}{\text{O}}_2(\text{g}) + \overset{+4}{\text{S}}\overset{-2}{\text{O}}_2(\text{g})</math>                      The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction.</p>
<p><b>Step 2</b> Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a decrease in oxidation number.</p>	<p>Each of two sulfur atoms lost 6 electrons and each of two oxygen atoms gained 2 electrons. So the <math>\text{O}_2</math> lost 12 electrons and the <math>\text{CS}_2</math> gained 4 electrons.</p> $\begin{array}{c} \text{1}(-12\text{e}^-) = -12\text{e}^- \\ \downarrow \\ \overset{+4}{\text{C}}\overset{-2}{\text{S}}_2(\text{g}) + \overset{0}{\text{O}}_2(\text{g}) \rightarrow \overset{+4}{\text{C}}\overset{-2}{\text{O}}_2(\text{g}) + \overset{+4}{\text{S}}\overset{-2}{\text{O}}_2(\text{g}) \\ \uparrow \qquad \qquad \qquad \uparrow \\ \text{3}(+4\text{e}^-) = +12\text{e}^- \end{array}$

# Chemistry Solutions Manual Redox Reactions

**Wenbin Ji**



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Nuclear Science Abstracts ,1974 Chemical Principles Monty J. Montjar,1990

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### **Table of Contents Chemistry Solutions Manual Redox Reactions**

1. Understanding the eBook Chemistry Solutions Manual Redox Reactions
  - The Rise of Digital Reading Chemistry Solutions Manual Redox Reactions
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Solutions Manual Redox Reactions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Solutions Manual Redox Reactions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Solutions Manual Redox Reactions
  - Personalized Recommendations

- Chemistry Solutions Manual Redox Reactions User Reviews and Ratings
- Chemistry Solutions Manual Redox Reactions and Bestseller Lists
- 5. Accessing Chemistry Solutions Manual Redox Reactions Free and Paid eBooks
  - Chemistry Solutions Manual Redox Reactions Public Domain eBooks
  - Chemistry Solutions Manual Redox Reactions eBook Subscription Services
  - Chemistry Solutions Manual Redox Reactions Budget-Friendly Options
- 6. Navigating Chemistry Solutions Manual Redox Reactions eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Solutions Manual Redox Reactions Compatibility with Devices
  - Chemistry Solutions Manual Redox Reactions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Solutions Manual Redox Reactions
  - Highlighting and Note-Taking Chemistry Solutions Manual Redox Reactions
  - Interactive Elements Chemistry Solutions Manual Redox Reactions
- 8. Staying Engaged with Chemistry Solutions Manual Redox Reactions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Solutions Manual Redox Reactions
- 9. Balancing eBooks and Physical Books Chemistry Solutions Manual Redox Reactions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Solutions Manual Redox Reactions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Solutions Manual Redox Reactions
  - Setting Reading Goals Chemistry Solutions Manual Redox Reactions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Solutions Manual Redox Reactions
  - Fact-Checking eBook Content of Chemistry Solutions Manual Redox Reactions

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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