

# Chemistry Review Packet

Name: \_\_\_\_\_

Date: \_\_\_\_\_

1. Consider the following reaction:



When 2.0 moles of hydrogen are reacted with 48.0 grams of oxygen, how many grams of water are formed?

- A. 18.0 grams    B. 36.0 grams    C. 54.0 grams    D. 11.0 grams    E. 90.0 grams

2. Which of the following 1.0 M aqueous solutions has the lowest pH?

- A.  $\text{HNO}_3$     B.  $\text{NaCl}$     C.  $\text{HNO}_2$     D.  $\text{NaOH}$     E.  $\text{HCO}_2\text{H}$

3. Which of the following substances cannot be broken down into simpler substances by physical or chemical means?

- A. Wood    B. Baking soda    C. Natural gas    D. Neon gas    E. Water

4. What is the approximate percent composition by mass of hydrogen in the compound ethanol,  $\text{C}_2\text{H}_6\text{O}$  (molecular weight, in grams: 46.10 or 46.08 or 46.0)?

- A. 4.3%    B. 12.0%    C. 29.0%    D. 34.0%    E. 20.0%

5. The empirical formula of a certain hydrocarbon is  $\text{C}_2\text{H}_4$ , and it has a molecular weight equal to 80 g/mol. Based on this information, the molecular formula of the hydrocarbon is

- A.  $\text{C}_2\text{H}_4$     B.  $\text{C}_4\text{H}_8$     C.  $\text{C}_6\text{H}_{12}$     D.  $\text{C}_8\text{H}_{16}$     E.  $\text{C}_{10}\text{H}_{20}$

6. Which of the following aqueous solutions *all* have the highest normal boiling point?

- A. 0.20 M  $\text{CaCl}_2$     B. 0.20 M  $\text{NaCl}$     C. 0.20 M  $\text{KNO}_3$     D. 1.0 M  $\text{C}_2\text{H}_6\text{O}_2$     E. 0.20 M  $\text{MgCl}_2$

7. Consider the reaction



How is this reaction best classified?

- A. Precipitation    B. Redox reaction    C. Single replacement    D. Combustion    E. Synthesis

8. The following reaction takes place under constant temperature and pressure:



When 10.0 L of hydrogen gas reacts completely with excess  $\text{H}_2$  again, how many liters of gas would form?

- A. 1.00 L    B. 1.50 L    C. 10.0 L    D. 10.0 L    E. 10.0 L

# Chemistry Slo Review Packet

**RS Peters**



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## **Table of Contents Chemistry Slo Review Packet**

1. Understanding the eBook Chemistry Slo Review Packet
  - The Rise of Digital Reading Chemistry Slo Review Packet
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Slo Review Packet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Slo Review Packet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Slo Review Packet
  - Personalized Recommendations
  - Chemistry Slo Review Packet User Reviews and Ratings
  - Chemistry Slo Review Packet and Bestseller Lists
5. Accessing Chemistry Slo Review Packet Free and Paid eBooks

- Chemistry Slo Review Packet Public Domain eBooks
  - Chemistry Slo Review Packet eBook Subscription Services
  - Chemistry Slo Review Packet Budget-Friendly Options
6. Navigating Chemistry Slo Review Packet eBook Formats
- ePub, PDF, MOBI, and More
  - Chemistry Slo Review Packet Compatibility with Devices
  - Chemistry Slo Review Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemistry Slo Review Packet
  - Highlighting and Note-Taking Chemistry Slo Review Packet
  - Interactive Elements Chemistry Slo Review Packet
8. Staying Engaged with Chemistry Slo Review Packet
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Slo Review Packet
9. Balancing eBooks and Physical Books Chemistry Slo Review Packet
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Slo Review Packet
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Slo Review Packet
- Setting Reading Goals Chemistry Slo Review Packet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Slo Review Packet
- Fact-Checking eBook Content of Chemistry Slo Review Packet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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