

Percent Composition Practice 2.1 – Answer Key

Objective: - Calculate the percent composition of a compound.

Directions: Determine the percent composition of each of the compounds below. (Show all work)

1. NaCl

$$\text{Na: } 22.99\text{g} \times 2 = 45.98\text{g}$$

$$\left(\frac{45.98\text{g}}{154.04\text{g}} \right) = 29.84\%$$

$$\text{Cl: } 35.45\text{g} \times 1 = 35.45\text{g}$$

$$\left(\frac{35.45\text{g}}{154.04\text{g}} \right) = 22.99\%$$

$$100.00\%$$

2. water

$$\text{H: } 1.00\text{g} \times 2 = 2.00\text{g}$$

$$\left(\frac{2.00\text{g}}{18.02\text{g}} \right) = 100\% = 11.11\%$$

$$\text{O: } 16.00\text{g} \times 1 = 16.00\text{g}$$

$$\left(\frac{16.00\text{g}}{18.02\text{g}} \right) = 100\% = 88.89\%$$

$$100.00\%$$

3. $\text{Cu}(\text{NO}_3)_2$

$$\text{Cu: } 63.55\text{g} \times 1 = 63.55\text{g}$$

$$\left(\frac{63.55\text{g}}{277.69\text{g}} \right) = 22.88\%$$

$$\text{N: } 14.01\text{g} \times 2 = 28.02\text{g}$$

$$\left(\frac{28.02\text{g}}{277.69\text{g}} \right) = 10.09\%$$

$$\text{O: } 16.00\text{g} \times 6 = 96.00\text{g}$$

$$\left(\frac{96.00\text{g}}{277.69\text{g}} \right) = 34.59\%$$

$$100.57\%$$

4. $\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_4$

$$\text{Fe: } 55.85\text{g} \times 1 = 55.85\text{g}$$

$$\left(\frac{55.85\text{g}}{274.05\text{g}} \right) = 20.38\%$$

$$\text{C: } 12.01\text{g} \times 8 = 96.08\text{g}$$

$$\left(\frac{96.08\text{g}}{274.05\text{g}} \right) = 34.91\%$$

$$\text{H: } 1.00\text{g} \times 6 = 6.00\text{g}$$

$$\left(\frac{6.00\text{g}}{274.05\text{g}} \right) = 2.19\%$$

$$\text{O: } 16.00\text{g} \times 8 = 128.00\text{g}$$

$$\left(\frac{128.00\text{g}}{274.05\text{g}} \right) = 46.72\%$$

$$100.00\%$$

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SJ Ball



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