

- Arrhenius Definitions:** When reacting with water
- Acids release hydrogen ions (ex: HNO<sub>3</sub>)
  - Bases release hydroxide ions (ex: NaOH)
  - Salts are ionic compounds that release neither hydrogen ions nor hydroxide ions

Practice: Write "A" for acid, "B" for base, and "S" for salt.

A
B
S  
 \_\_\_\_\_  
B
A
 \_\_\_\_\_  
 \_\_\_\_\_

Solution	Acid (A), Base (B), or Salt (S)	Reacts with Mg (R or NR)	Conductivity (C or NC)
HNO <sub>3</sub>	A	R	C
NaOH	B	NR	C
NaCl	S	NR	NC
H <sub>2</sub> SO <sub>4</sub>	A	R	C
HCl	A	R	C
NH <sub>4</sub> OH	B	NR	C
KNO <sub>3</sub>	S	NR	NC
KOH	B	NR	C
HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	A	R	C

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**Joacim Rocklöv**



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