

Name _____ Date _____ Class _____

CHAPTER 3 STUDY GUIDE FOR CONTENT MASTERY

Matter—Properties and Changes

Section 3.1 Properties of Matter

In your textbook, read about physical properties and chemical properties of matter.

Use each of the terms below just once to complete the passage.

chemical density	mass properties	physical substance
---------------------	--------------------	-----------------------

Matter is anything with (1) mass and volume. A(2) substance is a form of matter with a uniform and unchanging composition.Substances have specific, unchanging (3) properties that can be observed.Substances have both physical and chemical properties. (4) Physical

properties can be observed without changing a substance's chemical composition. Color,

hardness, and (5) density are examples. Other properties cannot be

observed without changing the composition of a substance. These are called

(6) chemical properties. An example is the tendency of iron to form

rust when exposed to air.

Label each property as either physical or chemical.

- physical 7. Chemical formula H_2O
- chemical 8. Irons gain cohesion when exposed to heat or
- chemical 9. Remains unchanged when in the presence of nitrogen
- physical 10. Crystals
- physical 11. Solid at normal temperatures and pressures
- chemical 12. Ability to combine with another substance
- physical 13. Melting point
- physical 14. Liquid at normal temperatures and pressures
- physical 15. Boiling point is $100^\circ C$
- physical 16. Conducts electricity
- physical 17. Density is $\frac{1}{8} \frac{g}{cm^3}$

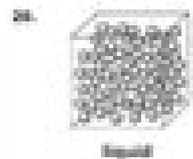
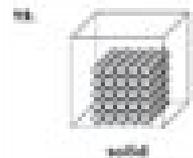
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CHAPTER 3 STUDY GUIDE FOR CONTENT MASTERY

Section 3.1 Matter

In your textbook, read about states of matter.

Label each drawing with one of these words: solid, liquid, gas.

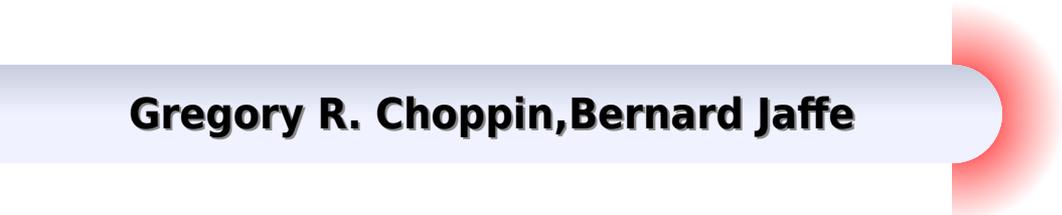


For each statement below, write true or false.

- true 21. All matter that we encounter in everyday life exists in one of three physical states.
- true 22. A solid has definite shape and volume.
- false 23. A liquid has a definite shape and takes on the volume of its container.
- true 24. A gas has both the shape and the volume of its container.
- false 25. The particles in a gas cannot be compressed into a smaller volume.
- false 26. Liquids tend to contract when heated.
- false 27. The particles in a solid are spaced far apart.
- false 28. The words gas and vapor can be used interchangeably.

Chemistry Matter Change Chapter 12 Answers

Gregory R. Choppin, Bernard Jaffe



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Solutions Manual for Chemistry: Molecules Matter and Change, Fourth Edition Julie Henderleiter, Charles Trapp, Peter Atkins, Loretta Jones, 1999-08-13 This student companion is a supplement to Chemistry Molecules Matter and Change 4th edition with CD ROM It features guided reading strategies collaborative learning sheets and strategies for using CD ROM tools

Chemistry: Molecules, Matter, and Change Media Activities Book Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change

Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change

Introduction to General, Organic, and Biological Chemistry Matta, Michael S. Matta, Dennis D. Staley, Anthony C. Wilbraham, 1996

Chemical Control of Gas Manufacture, 1916 *Chemistry: Science of Matter, Energy, and Change* Gregory R. Choppin, Bernard Jaffe, 1965

Essentials of Chemical Education Hans-Dieter Barke, Günther Harsch, Siegbert Schmid, 2011-10-28 For everybody teaching chemistry or becoming a chemistry teacher the authors provide a practice oriented overview with numerous examples from current chemical education including experiments models and exercises as well as relevant results from research on learning and teaching With their proven concept the authors cover classical topics of chemical education as well as modern topics such as every day life chemistry student s misconceptions the use of media or the challenges of motivation This is the completely revised and updated English edition of a highly successful German title

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CHEMISTRY SILBERBERG, 2003

Student Solutions Manual to Accompany Chemistry William Joseph Pietro, 1990

Matter and Interactions Ruth W. Chabay, Bruce A. Sherwood, 2015-01-12 Matter and Interactions 4th Edition offers a modern curriculum for introductory physics calculus based It presents physics the way practicing physicists view their discipline while integrating 20th Century physics and computational physics The text emphasizes the small number of fundamental principles that underlie the behavior of matter and models that can explain and predict a wide variety of physical phenomena Matter and Interactions 4th Edition will be available as a single volume hardcover text and also two paperback volumes

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