

CHAPTER 3 CHAPTER ASSESSMENT

MATTER AND CHANGE

Reviewing Vocabulary

Match the definition in Column A with the term in Column B.

Column A

- _____ 1. The outermost electrons of an atom
- _____ 2. The center of an atom
- _____ 3. Hot, highly ionized, electrically conducting gas
- _____ 4. The smallest particle of an element that retains that element's characteristics
- _____ 5. The change from a gas to a liquid
- _____ 6. A homogeneous mixture
- _____ 7. The change of one or more substances into other substances
- _____ 8. The attractive force between two ions of opposite charge
- _____ 9. A tiny particle with a negative electrical charge
- _____ 10. The average of the mass numbers of an element's isotopes
- _____ 11. The spontaneous process through which unstable nuclei emit radiation

Column B

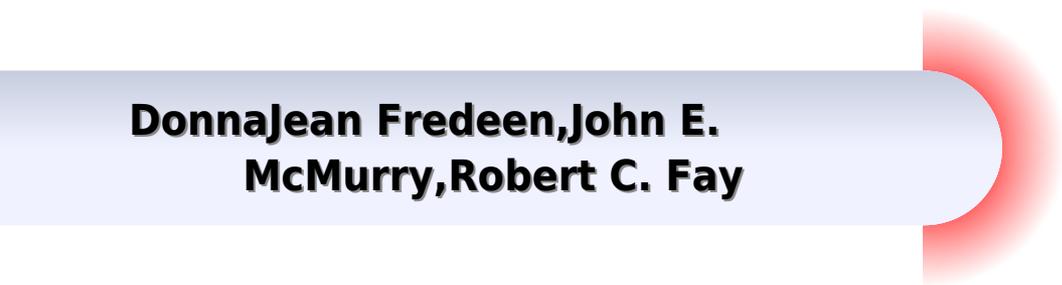
- a. atom
- b. atomic mass
- c. chemical reaction
- d. condensation
- e. electron
- f. ionic bond
- g. solution
- h. nucleus
- i. plasma
- j. valence electrons
- k. radioactive decay

In the space at the left, write the word or phrase in parentheses that makes the statement correct.

- _____ 12. The combined number of protons and neutrons is the (atomic number, mass number).
- _____ 13. An (energy level, isotope) represents the area in an atom where an electron is most likely to be found.
- _____ 14. (An ionic bond, A covalent bond) is the attraction of two atoms for a shared pair of electrons that holds the atoms together.
- _____ 15. An atom that gains or loses an electron is a charged particle called (an ion, a molecule).
- _____ 16. (An element, A compound) is a substance that cannot be broken down into simpler substances.
- _____ 17. A tiny particle that has mass and a positive electrical charge is a (neutron, proton).
- _____ 18. (Acids, Bases) are solutions characterized by the formation of hydroxide ions (OH^-).
- _____ 19. The process of changing from a liquid into a gas is called (evaporation, sublimation).

Chemistry Matter And Change Chapter 10 Assessment Answer Key

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McMurry, Robert C. Fay**



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Books in Print Supplement, 2002 **Chemistry** John S. Phillips, Cheryl Wistrom, 2000 **Holt Chemistry** Ralph Thomas Myers, 2004 Principles of Science Charles H. Heimler, 1983 Resources for Teaching Middle School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-03-30

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