

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



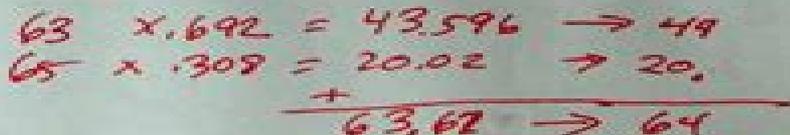
Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



Chemistry Guided Question Answers

Guillaume Favre

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Table of Contents Chemistry Guided Question Answers

1. Understanding the eBook Chemistry Guided Question Answers
 - The Rise of Digital Reading Chemistry Guided Question Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Guided Question Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Guided Question Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Guided Question Answers
 - Personalized Recommendations
 - Chemistry Guided Question Answers User Reviews and Ratings
 - Chemistry Guided Question Answers and Bestseller Lists
5. Accessing Chemistry Guided Question Answers Free and Paid eBooks

- Chemistry Guided Question Answers Public Domain eBooks
 - Chemistry Guided Question Answers eBook Subscription Services
 - Chemistry Guided Question Answers Budget-Friendly Options
6. Navigating Chemistry Guided Question Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Guided Question Answers Compatibility with Devices
 - Chemistry Guided Question Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Guided Question Answers
 - Highlighting and Note-Taking Chemistry Guided Question Answers
 - Interactive Elements Chemistry Guided Question Answers
 8. Staying Engaged with Chemistry Guided Question Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Guided Question Answers
 9. Balancing eBooks and Physical Books Chemistry Guided Question Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Guided Question Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Guided Question Answers
 - Setting Reading Goals Chemistry Guided Question Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Guided Question Answers
 - Fact-Checking eBook Content of Chemistry Guided Question Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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