

Chemical Equilibrium**STUDY QUESTIONS & PROBLEMS**

1. Write the expressions for the equilibrium constant
- K_c
- for the following reactions:



2. Write the equilibrium constant expressions for the following reactions. How are they related to one another?



3. Calculate the value of the equilibrium constant for the following system, given the data shown:



Concentrations at equilibrium:

$$[\text{H}_2] = 1.5 \text{ mol Liter}^{-1}$$

$$[\text{CO}_2] = 2.5 \text{ mol Liter}^{-1}$$

$$[\text{H}_2\text{O}] = 0.5 \text{ mol Liter}^{-1}$$

$$[\text{CO}] = 3.0 \text{ mol Liter}^{-1}$$

4. Chlorine molecules will dissociate at high temperatures into chlorine atoms. At 3000°C, for example,
- K_c
- for the equilibrium shown is 0.55. If the partial pressure of chlorine molecules is 1.5 atm, calculate the partial pressure of the chlorine atoms:
-
- $$\text{Cl}_2(\text{g}) \rightleftharpoons 2 \text{Cl}(\text{g})$$

5. Suppose that 0.50 moles of hydrogen gas, 0.50 moles of iodine gas, and 0.75 moles of hydrogen iodide gas are introduced into a 2.0 Liter vessel and the system is allowed to reach equilibrium.

Calculate the concentrations of all three substances at equilibrium. At the temperature of the experiment, K_c equals 2.0×10^{-2} .

6. If the mechanism of a chemical equilibrium consists of two reversible elementary steps, each with its own equilibrium constant
- K_{c1}
- and
- K_{c2}
- , what expression relates the equilibrium constant
- K_c
- for the overall equilibrium to the two constants
- K_{c1}
- and
- K_{c2}
- ?

7. When 2.0 mol of carbon disulfide and 4.0 mol of chlorine are placed in a 1.0 Liter flask, the following equilibrium system results. At equilibrium, the flask is found to contain 0.30 mol of carbon tetrachloride. What quantities of the other components are present in this equilibrium mixture?



Chemistry Equilibrium Practice Problems Calculation

Robert R. Redfield



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

Ebook: Chemistry Julia Burdge, 2014-10-16 Chemistry Third Edition by Julia Burdge offers a clear writing style written with the students in mind. Julia uses her background of teaching hundreds of general chemistry students per year and creates content to offer more detailed explanation on areas where she knows they have problems. With outstanding art, a consistent problem solving approach, interesting applications woven throughout the chapters, and a wide range of end of chapter problems, this is a great third edition text.

Chemistry for the IB Diploma Geoff Neuss, 2001 This concise guide provides the content needed for the Chemistry IB diploma at both Standard and Higher Level. It follows the structure of the IB Programme exactly and includes all the options. Each topic is presented on its own page for clarity. Higher Level material is clearly indicated and there are plenty of practice questions. The text is written with an awareness that English might not be the reader's first language.

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web continuing applicants years 2 10 for years two through ten submit only the following documents with required signatures by april 15 for the following school year lausd

division of risk management and insurance services fmla - Nov 22 2021

classified staff resignation form 5209 los angeles unified - Nov 03 2022

web the leave of absence loa scheme applies to all singapore citizen sc and permanent resident pr children studying in primary schools secondary schools junior colleges

los angeles unified school district schoolwires - Jul 31 2022

web may 23 2023 lausd retention form from date principal the los angeles unified school district you will receive an email confirming submission of this form this bulletin

pcss school forms and resources los angeles unified - Jul 11 2023

web retention a parent who agrees to have their child continue in kindergarten must sign the kindergarten continuance form for more information refer to ref 6756 1

local district superintendents community of schools number - Jun 10 2023

web 2016 lausd login bulletins no resource specified bul 6825 0 records retention and destruction other bul 6825 0 records retention and destruction other than pupil

los angeles unified school district memorandum - Jan 05 2023

web enter your full lausd email address and password to log in e g msmith lausd net mary smith lausd net