

Name _____

Electrons in Atoms



Study Guide: Properties of Waves and Quantum Mechanics

Properties of Waves

- **Wavelength (λ):** The distance between successive crests of a wave, usually measured in meters.
- **Frequency (ν):** The number of waves that pass a given point per second, measured in Hertz (Hz).
- **Amplitude:** The height of the wave crest or depth of the trough, related to the wave's energy.
- **Speed of light (c):** The rate at which the wave propagates through a medium, calculated using $c = \lambda\nu$

Calculating the Wavelength of a Wave

$$\lambda = c/\nu$$

- λ is the wavelength.
- c is the wave speed.
- ν is the frequency.

Speed of Light

- The speed of light in a vacuum (c) is approximately 3.00×10^8 meters per second (m/s).

The Quantum Equation

- Relates energy of a photon to its frequency: $E = h\nu$
 - E is the energy.
 - h is Planck's constant (6.626×10^{-34} Js).
 - ν is the frequency.

Element Emission Spectrum

The concept of quantized energy levels within atoms is crucial to understanding atomic spectra. When an electron absorbs energy, it jumps to a higher energy level, and when it returns to a lower energy level, it emits a photon of light with a specific energy corresponding to the difference in energy levels. This explains why each element has a unique spectral fingerprint, a set of specific wavelengths of light it emits or absorbs.

- When elements are heated, they emit light at specific wavelengths, creating an emission spectrum. This spectrum is unique to each element and can be used to identify them.



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