

The Chemistry of Cookies

Get this "bake" gone with the chemistry lab. Students investigate the role of each ingredient in the production of the cookies. Students will learn about the chemical reactions that occur during the baking process and the role of each ingredient in the production of the cookies. This will be a fun and educational activity for students.

Table 1: Molar Masses of Ingredients (Use the following molar masses to calculate the mass of ingredients needed to make 1 dozen cookies. Assume that the ingredients are 100% pure.)

Butter	$C_{18}H_{36}O_2$	Flour	$C_6H_{10}O_5$
Baking Soda	$NaHCO_3$	Egg	$C_5H_9NO_6$
Baking Powder	$NaHCO_3$	Egg Yolk	$C_5H_9NO_6$
Brown Sugar	$C_{12}H_{22}O_{11}$	Egg White	$C_5H_9NO_6$
Salt	$NaCl$	Vanilla	$C_{10}H_{12}O_2$
Vanilla Sugar	$C_{12}H_{22}O_{11}$	Vanillin	$C_{10}H_{12}O_2$
Chocolate Chips	$C_{15}H_{18}O_2$	Vanillin	$C_{10}H_{12}O_2$
Confectioner's Sugar	$C_{12}H_{22}O_{11}$	Vanillin	$C_{10}H_{12}O_2$
Sugar	$C_{12}H_{22}O_{11}$	Vanillin	$C_{10}H_{12}O_2$

Table 2: Recipe Ingredients (Use the following ingredients to make 1 dozen cookies. Assume that the ingredients are 100% pure.)

- 1 cup of butter = 227 g
- 1 cup of baking soda = 250 g
- 1 cup of vanilla extract = 237 g
- 1 cup of sugar = 200 g
- 1 cup of flour = 125 g
- 1 cup of confectioner's sugar = 125 g
- 1 cup of salt = 250 g
- 1 cup of eggs = 100 g
- 1 cup of vanilla = 100 g
- 1 cup of chocolate chips = 100 g

Ingredients

1 cup of butter (227 g) (100% pure)
 1 cup of baking soda (250 g) (100% pure)
 1 cup of vanilla extract (237 g) (100% pure)
 1 cup of sugar (200 g) (100% pure)
 1 cup of flour (125 g) (100% pure)
 1 cup of confectioner's sugar (125 g) (100% pure)
 1 cup of salt (250 g) (100% pure)
 1 cup of eggs (100 g) (100% pure)
 1 cup of vanilla (100 g) (100% pure)
 1 cup of chocolate chips (100 g) (100% pure)



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Yicheng Fang



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