

Chemical Bonding

GROUP STUDY GUIDE

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1. Bonding & Stability

Atoms form bonds to:

- Decrease potential energy
- Increase stability

Why do atoms form bonds?

Potential energy is decreased as bond is formed.

Force of attraction

Force of repulsion

The forces of attraction balance at a distance where potential energy is at a minimum, and a bond is formed.

Atomic ions & ionic compounds

Polytatomic ions

NO₃⁻
 $8e^- + 5 + 3(6) = 26 + 18 = 44e^-$

NH₄⁺
 $5e^- + 4(1) = 9 + 4 = 13e^-$

When chemical bonds form, valence electrons are transferred from one atom to another, shared between atoms, or delocalized in a "sea" of electrons, metals.

The electronegativity difference between two bonded atoms determines the type of bond.

- 0.0 - 0.3 = non-polar covalent
- 0.3 - 1.7 = polar covalent
- 1.7 - and greater = ionic

Two major categories of compounds based on bonding are ionic and molecular (covalent) compounds.

ionic compounds: metal bonded with a nonmetal, or a compound containing a polyatomic ion.

molecular (covalent) compounds: two or more nonmetals combined.

Bonding guidelines:

- Metals + nonmetals = ionic compound
- Metals + nonmetals + covalent compound (molecular)
- Nonmetal + nonmetal = covalent compound
- Nonmetal + nonmetal + polyatomic ions = ionic compound

Step 2: Draw the skeletal structure.

Step 5: Count the electrons.

24e⁻

CHEMISTRY CORNER

Chemistry Chemical Bonding Study Guide

RS Peters



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