

# Chemistry 105

## Chapter 3 Review Problems

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- How many molecules are present in 180 grams of water? (The atomic weights of H and O are 1 g/mol and 16 g/mol, respectively.)
  - $6.0 \times 10^{24}$
  - 10
  - $5.6 \times 10^{-23}$
  - $1.8 \times 10^7$
  - $180 \times 10^{23}$
- The atomic mass of sulfur is given as 32.07 g/mol. What is the mass of an average atom of sulfur?
  - $5.33 \times 10^{-23}$  g
  - 32.07 g
  - $32.07 \times 10^{-23}$  kg
  - $193 \times 10^{-23}$  g
  - $6.02 \times 10^{23}$  g
- What is the mass of C present in 60 g of  $C_3H_6$ ?
  - 12 g
  - 24 g
  - 48 g
  - 60 g
  - 6 g
- What is the molar mass of  $C_3H_6O$ ?
  - 24 g/mol
  - 32 g/mol
  - 7 g/mol
  - 35 g/mol
  - 46 g/mol
- What is the correct percentage by mass of nitrogen in HCN?
  - 33 %
  - 14%
  - 27%
  - 52%
  - 98%
- What is the chief danger that is anticipated if  $CO_2$  levels get too high in the atmosphere?
  - people will develop skin cancer from exposure to UV light
  - the atmosphere will develop an unpleasant odor
  - the temperature of the planet may climb by several degrees
  - plant life will be exterminated
  - The  $CO_2$  will remove ozone from the stratosphere

# Chemistry Chapter 3 Review Answers

**Susan McMurry**



## Chemistry Chapter 3 Review Answers:

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1. Open the passenger door and remove the dash trim on the side of the dash. (Figure A).
2. Open the glove box and remove.