

Chapter 9 Review

- 1) What does stoichiometry mean? "measuring elements"
- 2) Before performing a stoichiometry problem, you must first do what with the chemical equation? **balance it**
- 3) The coefficients in a chemical equation represent what? **moles**
- 4) In a chemical equation, what relationships are shown? **molar ratios**
- 5) Solve the following mole-mole conversions using this equation:



- a) If 3.5 moles of sodium bicarbonate are decomposed, how many moles of carbon dioxide are produced?

$$3.5 \text{ mol NaHCO}_3 \times \frac{1 \text{ mol CO}_2}{2 \text{ mol NaHCO}_3} = 1.8 \text{ mol CO}_2$$

- b) If 1.29 moles of water are produced by the decomposition of sodium bicarbonate, how many moles of sodium carbonate are produced?

$$1.29 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{1 \text{ mol H}_2\text{O}} = 1.29 \text{ mol Na}_2\text{CO}_3$$

- c) How many moles of sodium bicarbonate are required to produce 19.5 moles of carbon dioxide?

$$19.5 \text{ mol CO}_2 \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol CO}_2} = 39.0 \text{ mol NaHCO}_3$$

- 6) Use the equation in # 5 to solve the following mass-mass conversions.

- a) If given 10.5 grams of sodium bicarbonate, how many grams of sodium carbonate can be produced?

$$10.5 \text{ g NaHCO}_3 \times \frac{1 \text{ mol NaHCO}_3}{84.01 \text{ g NaHCO}_3} \times \frac{1 \text{ mol Na}_2\text{CO}_3}{2 \text{ mol NaHCO}_3} \times \frac{105.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} = 6.62 \text{ g Na}_2\text{CO}_3$$

- b) How many grams of sodium bicarbonate would be required to produce 9.0 grams of water?

$$9.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol NaHCO}_3}{1 \text{ mol H}_2\text{O}} \times \frac{84.01 \text{ g NaHCO}_3}{1 \text{ mol NaHCO}_3} = 84 \text{ g NaHCO}_3$$

- 7) What conversion factor (mole ratio) would you use to solve for moles of sulfur formed from sulfur dioxide in the reaction



Chemistry Chapter 1stoichiometry Packet Answers

Clyde Metz



Chemistry Chapter 1 stoichiometry Packet Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03

The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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