

General Chemistry II Chapter 17 Study Guide

Organic acids- carboxylic acids (R-COOH):

- Carboxylic acids (R-COOH): a weak organic acid, contains the carb-oxyl group (-COOH), never fully dissociates in water, an equilibrium process, ex. Formic acid, acetic acid, butyric acid, benzoic acid
 - R= rest of the molecule
- Strength of carboxylic acids: depends on the R group; more electronegative Cl in R group shifts more electron density towards the R group and helps stabilize the structure after the H⁺ ion is removed, thus, chloroacetic acid is the stronger acid in comparison to acetic acid (CH₃COOH)

Organic bases- amines:

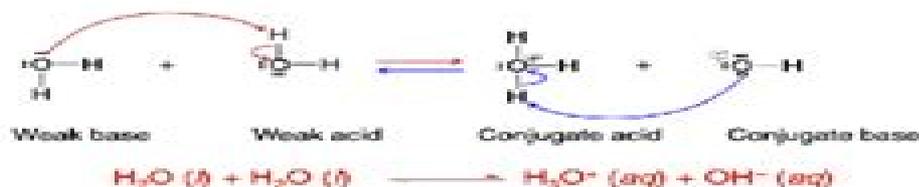
- Amines: organic bases, contains an N atom with a lone pair (ex -NH, -NH₂), ex. Methyl amine, dimethyl amine, trimethyl amine; are weak bases:
 - Only N with lone pair can accept a proton
 - Equilibrium dissociation process
 - NH₄⁺ is an acid and can lose one H⁺ to be neutral

Organic acids and bases (things to keep in mind):

- Acid: H is the only acidic proton; ex. CH₃COOH becomes CH₃COO⁻
- Amine: the proton H⁺ is bonded to N via the lone pair of N; ex. CH₃NH₂ becomes CH₃NH₃⁺
- How about carbon?
 - Protons on C are not acidic, C cannot accept extra protons

Autoionization of water:

- Autoionization: in pure water, one water molecule can donate a proton to another water molecule



At 25 °C, [H₃O⁺] = [OH⁻] = 1.0 x 10⁻⁷ M

$$K_w = [\text{H}_3\text{O}^+] [\text{OH}^-] = 1.0 \times 10^{-14}$$

K_w = ion-product constant

Chemistry Chapter 17 Study Guide Answers

Yan Bai



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Table of Contents Chemistry Chapter 17 Study Guide Answers

1. Understanding the eBook Chemistry Chapter 17 Study Guide Answers
 - The Rise of Digital Reading Chemistry Chapter 17 Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 17 Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 17 Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 17 Study Guide Answers
 - Personalized Recommendations
 - Chemistry Chapter 17 Study Guide Answers User Reviews and Ratings
 - Chemistry Chapter 17 Study Guide Answers and Bestseller Lists

5. Accessing Chemistry Chapter 17 Study Guide Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Chapter 17 Study Guide Answers
 - Interactive Elements Chemistry Chapter 17 Study Guide Answers
8. Staying Engaged with Chemistry Chapter 17 Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 17 Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Chapter 17 Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 17 Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 17 Study Guide Answers
 - Setting Reading Goals Chemistry Chapter 17 Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 17 Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Chapter 17 Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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