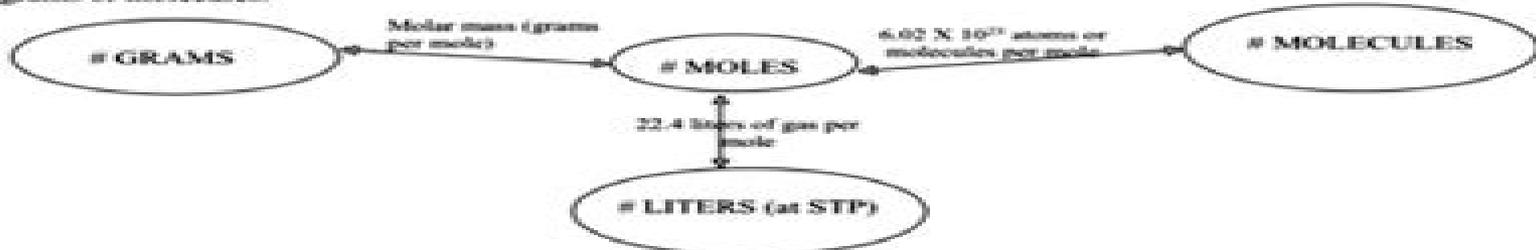


You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP. We can use this information to convert grams to molecules or liters, molecules to grams or liters, or liters to grams or molecules.



EXAMPLE 1: What would be the volume in liters of 40.36 grams of neon at STP?

$$\# \text{ liters Ne} = 40.36 \text{ g Ne} \times \frac{1 \text{ mole Ne}}{20.18 \text{ g Ne}} \times \frac{22.4 \text{ L Ne}}{1 \text{ mole Ne}} = 44.80 \text{ L Ne}$$

EXAMPLE 2: How many molecules would there be in 56 liters of carbon dioxide at STP?

$$\# \text{ molecules CO}_2 = 56.0 \text{ L CO}_2 \times \frac{1 \text{ mole CO}_2}{22.4 \text{ L CO}_2} \times \frac{6.02 \times 10^{23} \text{ molecules CO}_2}{1 \text{ mole CO}_2} = 1.51 \times 10^{24} \text{ molecules CO}_2$$

SOLVE THE FOLLOWING PROBLEMS ON A SEPARATE SHEET OF PAPER.

- YOU MUST USE COMPLETE AND PROPER SET-UPS.
- SHOW THE MOLAR MASS CALCULATION WHENEVER THE PROBLEM REQUIRES YOU TO DO ONE. INCLUDE UNITS
- CIRCLE YOUR FINAL ANSWER.

1. What would be the volume, in liters, of 85.5 grams of carbon monoxide at STP?
2. How many molecules would there be in 0.500 grams of carbon disulfide?
3. What would be the mass, in grams, of 45.0 liters of nitrogen at STP?
4. How many molecules of hydrogen are in a balloon full of hydrogen with a volume of 5.34 liters at STP?
5. Your mommy buys you a helium balloon at the circus. It has a volume of 4.00 liters at STP. What mass of helium, expressed in grams, does this balloon contain?
6. How many molecules of ammonia would there be in 40.0 grams of ammonia?
7. What would be the mass, in grams, of 3.50×10^{23} molecules of chlorine?
8. What volume, expressed in liters, would 50.0 grams of fluorine occupy at STP?
9. How many grams of oxygen would there be in 1.00 liter of oxygen at STP?

Extra Challenge:

10. How many molecules of water are there in 10 lbs of water?

Chemistry B Moles Packet Worksheet 6

Julie C. Gilbert



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