

AP Chemistry Study Guide for Ch. 17 – Additional Aspects of Aqueous Equilibria

Section 17.7 will not be tested.

Learning Goals	Exercises to Do
<p>17.1 Common Ion Effect Know how this is an example of LeChatelier's Principle. Common ion situation occurs when adding an ionic compound, a salt (MA), to a weak acid (HA) and the anion of the compound (A^-) is identical to the conjugate base of the weak acid (A^-). The anion (A^-) is the "common ion". This effect also occurs when adding a strong acid to a weak acid, in which the common ion is the proton, H^+.</p>	SE17.1/P1-2 SE17.2/P1-2 GIST p729 17.13, 17.15, 17.17
<p>17.2 Buffers Buffers make a solution resist changes in pH caused by addition of a small amount of strong acid or strong base. Study Fig. 17.2. Know the 2 ways to make a buffered solution – p. 729 Know how to predict pH of a buffered solution using the fast and exciting Henderson-Hasselbalch Equation. Know how to calculate pH of a buffered solution the slow ICE Table/Quadratic Equation Method. Study the comment at the end of SE 17.4 (middle of page 733) Know what determines the capacity of a buffered solution. Know how to form a buffer that has a particular pH value. Learn how to calculate the effect on pH when strong acid or strong base is added to a buffered solution.</p>	GIST p729 GIST p731 SE17.3/P1-2 SE17.4/P1-2 SE17.5/P1-2 SE17.6/P1-2 GIST p735 17.19, 17.21, 17.23, 17.25, 17.31
<p>17.3 Acid-Base Titrations The $pH = pK_a$ of a weak acid at the half-way point when titrating a weak acid with a strong base. The $pOH = pK_b$ of a weak base at the half-way point when titrating a weak base with a strong acid</p>	SE17.7/P1-2 GIST p.740 SE17.8/P1-2 SE17.9/P1-2 GIST p.744 GIST – both on p.746 17.5, 17.17 17.33, 17.35, 17.37, 17.39, 17.41, 17.45, 17.47
<p>17.4 Solubility Equilibria Know how to write a K_{sp} expression Know how to calculate K_{sp} Know how to determine solubility of a compound from its K_{sp}</p>	SE17.10/P1-2 Give it sum that p.749 SE17.11/P1-2 SE17.12/P1-2

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