

CHEMISTRY
Class – XII

Time – 3 hrs.

M.M. : 70

Date – 14.09.2022 (Wednesday)

Name of the student _____

Section _____

GENERAL INSTRUCTIONS:

- All questions are compulsory.
- Question nos. 1 to 21 are MCQs with one correct option and carrying 1 mark each.
- Question nos. 22 to 24 are objective/short answer type questions.
- Question nos. 25 to 28 are of 2 marks each.
- Question nos. 29 to 32 is of 3 marks each.
- Question nos. 33 to 35 is of 5 marks each.

SECTION - A

Q. 1 to 21 is of 1 mark each.

Q.1 Which of the following aqueous solution should have highest boiling point ?

- a) 1 M NaOH b) 1 M Na₂SO₄ c) 1M NH₄NO₃ d) 1 M KNO₃

Q.2 Which of the following statement is false ?

- a) Two different solution of sucrose of same molality prepare in different solvent will have the same depression in freezing point.
- b) The osmotic pressure of a solution is given by the equation $\pi = CRT$. (where C is the molarity of the solution)
- c) Decreasing order of osmotic pressure for 0.01M aqueous solution of Barium Chloride, Potassium chloride, acetic acid and sucrose is BaCl₂ > KCl > CH₃COOH > Sucrose
- d) According to Raoult's Law, the vapour pressure exeterd by a volatile component of solution is directly proportional to its molar fraction in the solution.

Q.3 We have three aqueous solution of NaCl labelled as 'A', 'B' and 'C' with concentration 0.1 M, 0.01 M and 0.001 M respectively. The value of Van't Hoff factor for these solution will be in the order-

- a) $i_A < i_B < i_C$ a) $i_A > i_B > i_C$ a) $i_A = i_B = i_C$ a) $i_A < i_B > i_C$

Q.4 On the basis of the information given below mark the correct options –

- i) In bromoethane and chloroethane mixture intermolecular interaction of A-A and B-B type are nearly same as A-B type interaction.
- ii) In the ethanol and acetone mixture A-A or B-B type intermolecular interactions are stronger than A-B type interaction.

Chemistry Question Paper For Class 12 Stateboard

P. W. Bedford



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CBSE Class 12 Mastering Organic Reactions MCQ With Answers Priti Singhal, 2024-11-11 The primary objective of this book is to serve as a comprehensive guide for students educators and researchers by focusing on reaction mechanisms practical applications and problem solving techniques Organic chemistry is not just about memorizing equations and formulas it is about understanding how molecules interact change and influence each other under different conditions With that in mind this book emphasizes the logic and patterns behind organic reactions making it easier for readers to apply concepts across a variety of scenarios Each chapter of this book builds upon foundational knowledge ensuring a progressive learning experience From nucleophilic substitutions to pericyclic reactions and from oxidation reduction mechanisms to named reactions we cover both fundamental and advanced topics to cater to students at all levels Real world examples have been integrated throughout the chapters to show how organic reactions play essential roles in pharmaceuticals biochemistry agriculture and environmental science This approach bridges the gap between theory and practical applications helping readers appreciate the relevance of organic chemistry in daily life

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