

Nuclear Chemistry - Guided Practice Questions**Multiple Choice** - Choose the one best answer and write its letter in the blank provided.

- ____ 1) In balancing the nuclear reaction ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th} + \text{_____}$, the identity of the missing element is _____?
 A) Pu B) U C) Th D) Np E) Pa
- ____ 2) The following rxn, ${}^{210}_{84}\text{Po} \rightarrow {}^{210}_{82}\text{Pb} + \text{_____}$, is an example of what form of radioactive decay?
 A) alpha B) positron C) gamma D) beta E) e-capture
- ____ 3) The missing product from this reaction is: ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th} + \text{_____}$?
 A) ${}^4_2\text{He}$ B) ${}^0_{-1}\text{e}$ C) ${}^4_2\text{He}$ D) ${}^0_1\text{n}$ E) ${}^0_{-1}\text{e}$
- ____ 4) What is the mass number of a neutron _____?
 A) 1 B) 3 C) 2 D) 0 E) 4
- ____ 5) In the nuclear decay reaction represented by ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th}$, the emitted particle is _____?
 A) beta particle B) neutron C) alpha particle D) positron E) proton
- ____ 6) The reaction shown below is responsible for creating C-14 in the atmosphere. What is the bombarding particle _____?
 ${}^{14}_7\text{N} + \text{_____} \rightarrow {}^{14}_6\text{C} + {}^1_1\text{H}$
 A) electron B) proton C) positron D) alpha particle E) neutron
- ____ 7) ${}^{131}_{53}\text{I}$ has a half-life of 8.04 days. Assuming you start with a 1.53 mg sample of ${}^{131}_{53}\text{I}$, how many mg will remain after 13.0 days _____?
 A) 0.440 B) 0.422 C) 0.298 D) 0.835 E) 0.499
- ____ 8) All atoms of a given element have the same:
 A) atomic mass, B) mass number, C) # of nucleons, D) atomic #, E) # of neutrons.
- ____ 9) Atoms containing radioactive nuclei are called:
 A) radioisotopes, B) radionuclides, C) nuclides, D) nucleons, E) radioisophores.
- ____ 10) What happens to the mass number and the atomic number of an element when it undergoes beta decay?
 A) The mass number decreases by 4 and the atomic number decreases by 2.
 B) The mass number does not change and the atomic number decreases by 2.
 C) Neither the mass number nor the atomic number change.
 D) The mass number increases by 2 and the atomic number increases by 1.
 E) The mass number does not change and the atomic number increases by 1.

Chemistry Guided Practice Problem Answer Key

Margaret G. Kimble



Chemistry Guided Practice Problem Answer Key:

GO TO Objective NEET 2021 Chemistry Guide 8th Edition Disha Experts, Study Guide with Selected Solutions Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems **Student Study Guide/solutions Manual to Accompany Foundations of Inorganic, Organic, and Biological Chemistry** Katherine J. Denniston, 1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface Invitation to Organic Chemistry Alyn William Johnson, 1999 Colorful graphics and 19 chapters featuring such learning aids as chemistry at work and conceptual problems characterize this large text on a large subject Cited by the American Association for the Advancement of Science for his pioneering work in the chemistry of ylides Johnson who spent most of his career at the U of North Dakota explores the smorgasbord of subject matter that is organic chemistry and new developments in the field Appends a summary of nomenclature spectra group assignments and values of selected important compounds The index is combined with a glossary Annotation copyrighted by Book News Inc Portland OR *Study Guide and Full Solutions Manual* Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook Learning to Solve Complex Scientific Problems David H. Jonassen, 2017-09-25 Problem solving is implicit in the very nature of all science and virtually all scientists are hired retained and rewarded for solving problems Although the need for skilled problem solvers has never been greater there is a growing disconnect between the need for problem solvers and the educational capacity to prepare them Learning to Solve Complex Scientific Problems is an immensely useful read offering the insights of cognitive scientists engineers and science educators who explain methods for helping students solve the complexities of everyday scientific problems Important features of this volume include discussions on how problems are represented by the problem solvers and how perception attention memory and various forms of reasoning impact the management of information and the search for solutions how academics have applied lessons from cognitive science to better prepare students to solve complex scientific problems gender issues in science and engineering classrooms and questions to guide future problem solving research The innovative methods explored in this practical volume will be of significant value to science and engineering educators and researchers as well as to instructional designers **Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition** John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

Introduction to General, Organic, and Biochemistry Morris Hein, Scott Pattison, Susan Arena, Leo R. Best, 2014-01-15

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Bongolan, Ellen Moir, Wendy Baron, 2009-11-16 The suggestions and tips give teachers a solid start to their year and help their students achieve This book has a unique focus on secondary teachers Janet Crews Secondary Instructional Coordinator Clayton School District MO A blueprint for successful teaching an entire year of teaching methods in one volume This book is jam packed with practical useful forms to guarantee a meaningful year for both the students and especially the teacher Ronald W Poplau Social Studies Teacher Shawnee Mission Northwest High School KS An easy to use source for all the strategies you need to thrive in the secondary classroom Preparing adolescent students for today's increasing demands requires commitment organization subject matter knowledge and passion But with the right strategies and tools first year secondary school teachers can approach this challenge with greater confidence and ease Leveraging a wealth of information and research from the New Teacher Center this user friendly guide provides a solid foundation for classroom management lesson planning and assessment Teachers will learn step by step tips for organizing standards based curriculum across the content areas supported by extensive reproducible forms and go to references including lesson plans from exemplary math and language arts teachers The authors provide specific strategies for Reaching struggling readers Working with English learners and their families Designing lessons with assessment in mind Sustaining a passion for teaching Covering everything needed for your first year and beyond this guide provides invaluable insights on succeeding with adolescent students

Essentials of Introductory Chemistry Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third

Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry

Student Study Guide Fundamentals of Chemistry by David E. Goldberg Margaret G. Kimble,1994

Introductory Chemistry Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students *Chemistry* John S. Phillips, Cheryl Wistrom, 2000 *Study Guide* Darrell D. Ebbing, Susan M. Schelble, 1998 Prentice Hall Chemistry , 2000

Organic and Biological Chemistry Karen Timberlake, 2001-11 Keyed to the learning goals in the Organic and Biological Chemistry text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems *Chemistry* George Tyler Miller, 1987

Academic Language/Literacy Strategies for Adolescents Debra L. Cook Hirai, Irene Borrego, Emilio Garza, Carl Kloock, 2013-02-01 Fast paced practical and innovative this text for pre service and in service teachers features clear easily accessible lessons and professional development activities to improve the delivery of academic language literacy education across the content areas in junior middle school and high school classrooms Numerous hands on tools and techniques demonstrate the effectiveness of content area instruction for students in a wide variety of school settings particularly English language learners struggling readers and other special populations of students Based on a strong professional development model the authors have been instrumental in designing Academic Language Literacy Strategies for Adolescents addresses motivation attributes of academic language vocabulary theory and practice reading skills development grammar and writing A wealth of charts graphs and lesson plans give clear examples of academic language literacy strategies in action The appendices a key component of the practical applications developed in the text include a glossary exemplary lessons that address key content areas and a Grammar Handbook In this era of increased accountability coupled with rapid demographic change and challenges to traditional curricula and pedagogical methods educators will find this book to be a great resource

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