

## Chemistry - Chapter 7 Study Guide

### Section 7.1:

- *Valence electrons* – electrons in the outer energy level
  - responsible for bonding with other atoms
- How do you know how many valence electrons an element has?
  - Look at the group number
    - Group 1A = 1 valence electron
    - Group 7A = 7 valence electrons
    - Etc...
- *Electron Dot Diagrams* – a way of showing and keeping track of valence electrons
  - Written using the following steps:
    1. Write the element symbol – representing the nucleus and inner electrons
    2. Put one dot for each valence electrons,
    3. Don't forget to add them using Hund's rule – they don't pair up unless they have to.
- *The Octet Rule* – atoms tend to want to achieve a noble gas configuration because it is the most stable. Octet = eight electrons
- *Cations* – a positive ion (usually metals)
  - Lose electrons to attain a noble gas configuration
  - Written like  $\text{Na}^{1+}$
  - No dots are shown (on an electron dot diagram) because they lose the valence electrons they had
- *Anions* – a negative ion (usually non-metals)
  - Gain electrons to attain a noble gas configuration
  - Written like  $\text{P}^{3-}$
  - Dots are shown because valence electrons are gained
- 8 valence electrons = stable atom (noble gases.)

### Section 7.2:

- Anions (-) and cations (+) are held together by opposite charges; a bond is formed through the transfer of electrons to achieve a noble gas configuration (octet.)
- *Ionic Compounds:*
  - Also called salts
  - Made from a cation and an anion
    - Which means they're made of a metal and a non-metal
- **\*if you don't understand how to do ionic bonding, check the book or ask somebody to help... it's hard to explain\***
- The result of ionic bonding is an ionic compound, shown in a *formula unit*
  - Written like  $\text{Ca}_3\text{P}_2$

# Chemistry Chapter 7 Study Guide Answers

**James E. Brady**



## **Chemistry Chapter 7 Study Guide Answers:**

*Organic Chemistry Study Guide with Solutions Manual* Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer **Hazmat Chemistry Study Guide (Second Edition)** Jill Meryl Levy, 2011 *Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition* Susan McMurry, 2000 Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines **Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond's text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry **Study Guide, Chemical Principles, Fifth Edition, Zumdahl** Paul B. Kelter, Steven S. Zumdahl, 2004-04 *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers

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