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CHAPTER 9

STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 11 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	π bond
---------------	----------	------------	------------	------------

When sharing electron pairs, the attachment between atoms that results is called a(n) **covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **π bond**.

In your textbook, read about single and multiple bonds and bond strength.

Circle the letter of the choice that best completes the statement or answers the question.

- Is a pair from an element such as hydrogen, oxygen, and nitrogen normally covalent?
 - an ionic atom
 - an molecule containing three atoms
 - an molecule containing two atoms
 - an molecule containing four atoms
- How many electrons are shared in a double covalent bond?
 - one
 - two
 - three
 - four
- Bond length is the distance between
 - two molecules of the same substance
 - the nucleus of two attached atoms
 - the electron in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - As the electron bond, the strength of the bond increases.
 - As the electron bond, the number of the bond increases.
 - As the electron bond, the lower the electron is a.
 - As the electron bond, the lower the bond dissociation energy.

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CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 12 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

- false** 1. Binary molecular compounds are generally composed of a metal and a nonmetal.
- false** 2. The second element in the formula of a binary compound is named using the suffix -ite.
- false** 3. The prefix zero indicates three atoms.
- true** 4. The prefix four indicates six atoms.
- true** 5. In naming the first element in a formula, the prefix zero is not used.
- true** 6. For binary acids, the hydrogen part of the compound is named using the prefix hydro.
- false** 7. An oxyacid contains only two elements.
- false** 8. If the name of the cation of an oxyacid ends in -ite, the acid name contains the suffix -ate.

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- c 1. SO_2
- i 2. CO
- g 3. H_2CO_3
- f 4. H_2SO_4
- b 5. N_2O_5
- d 6. HNO_3
- i 7. HNO_2
- a 8. HBr
- h 9. HBrO_3

Column B

- a.** hydrobromic acid
- b.** nitrogen trioxide
- c.** carbon trioxide
- d.** nitrous acid
- e.** carbonic acid
- f.** nitric acid
- g.** carbonic acid
- h.** bromic acid
- i.** carbon dioxide

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