

ChemThink: Covalent Bonding

Consider bonding between two atoms with some distance between them. Electrons

When two atoms get closer together, the electrons attract the other atom's

Protons & Electrons.

Before bonding, the atoms's electrons spread over a large area

area. The closer the two atoms, (and bonded), the electrons spread over a smaller area. This less spread

electrons must be closer to both atoms than ever. Electrons will pull smaller atom's electrons.

Consider bonding between two electrons.

When atoms move closer, the potential energy increases & decreases.

At a certain point the potential energy decreases if you keep and lower the atoms closer together than point to which nuclear are repelling each other.

This ideal distance between the atoms is known as the bond length.

Lower in energy = more stable.

Bond Type	Example	# of shared e ⁻	Total # of e ⁻ shared	Strength
Single	H—H	1	2	weakest
Double	O=O	2	4	stronger
Triple	N≡N	3	6	strongest

Chemthink Covalent Bonding Answers

Leallyn Burr Clapp



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Chemistry of the Covalent Bond Leallyn Burr Clapp, 1957

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